Manipur University

Structure of

3 – Year B.Sc (Honours) Degree Course in

CHEMISTRY

Under Semester System

Total Marks : 1000

Duration : 3 Years

Examination	Theory	Practical	Total
1st Semester	75	25	100
2nd Semester	75	25	100
3rd Semester	75	25	100
4th Semester	75	25	100
5th Semester	200	100	300
6th Semester	200	100	300
Total	700	300	1000

Scheme of Academic Programme

Based on the assumption that there will be 180 working days in a year or 90working days in a Semester and there will be 6-days teaching in a week, the expected working days for effective teaching are 15 weeks per Semester.

The Schedule for Chemistry teaching:

Examination	Theory	Practical	Total Hrs
1st Semester	6 Hrs / Week 90 Hrs / Sem	3 Hrs / Week 45 Hrs / Sem	9Hrs / Week 135 Hrs / Sem
2nd Semester	do	do	do
3rd Semester	do	do	do
4th Semester	do	do	do
5th Semester	6 Hrs /Week per paper 18 Hrs/Week for Theory papers 90 Hrs per paper per semester 270 Hrs per Sem for Theory papers		
6th Semester	Do	do	

Question Pattern

Questions are to be set from every unit corresponding to the marks allocated for each unit; option, alterative questions if any may be given within the same unit not among the units.

Duration of Examination

Each Theory Paper: 3 Hours

a) Practical Papers CH-101P, CH-202P, CH-303P and CH-404P---6 Hours each
b) Practical Papers CH-508P and CH-611P---12Hours each spread over two days

Paper – Wise Marks Distribution

FirstSemester

Papers		Marks	Time (Hours)
CH-101	Section A :InorganicChemistry	25	30
	Section B :Organic Chemistry	25	30
	Section C :PhysicalChemistry	25	30
CH-101P	Practical, Inorganic Experiments	25	45

Second Semester

CH-202	Section A: Inorganic Chemistry	25	30
	Section B: Organic Chemistry	25	30
	Section C: Physical Chemistry	25	30
CH-202P	Practical, Organic Experiments	25	45

Third Semester

CH-303	Section A: Inorganic Chemistry	25	30
	Section B: Organic Chemistry	25	30
	Section C: Physical Chemistry	25	30
CH-303P	Practical, Physical Experiments	25	45

Fourth Semester

CH-404	Section A: Inorganic Chemistry	25	30
	Section B: Organic Chemistry	25	30
	Section C: Physical Chemistry	25	30
CH-404P	Practical, Analytical Experiments	25	45

Fifth Semester

CH-505	Inorganic Chemistry	67	90
CH-506	Organic Chemistry	67	90
CH-507	Physical Chemistry	66	90
CH-508P	a. Inorganic	67	135
	b. Physical	33	

Sixth Semester

CH-608	Inorganic Chemistry	67	90
CH-609	Organic Chemistry	66	90
CH-610	Physical Chemistry	67	90
CH-611P	a. Organic	67	135
	b. Physical	33	

SYLLABUS

BACHELOR OF SCIENCE

in

CHEMISTRY (HONS)

SEMESTER – I

CH-101

Section A : INORGANIC CHEMISTRY

25 Marks; 30 Hours

Unit 1 Atomic Structure 6 Marks

Idea of de Broglie Matter waves, Heisenberg uncertainty principle, atomic orbital's, Schrodinger wave equation, quantum numbers, radial and angular wave functions, and probability distribution curves, shapes of s, p, d, orbital's, Aufbau and Pauli exclusion principles, Hund's multiplicity rule, Electronic configurations of the elements, effective nuclear charge.

Unit 2 Periodic Classification of Elements 6 Marks

Electronic configuration of the elements, atomic and ionic radii, ionization energy, electron affinity, and electronegativity – definition methods of determination or evaluation, trends in periodic table and applications in predicting and explaining the chemical behavior.

Unit 3 Chemical Bonding 8 Marks

Covalent bond - Valence Bond theory and its limitations, directional characteristics of covalent bond, various types of hybridization and shapes of simple inorganic molecules and ions. Valence shell electron pair repulsion theory (VSEPRT) to NH_3, H_30^+, SF_4, CIF_3, ICI_2, Molecular orbital theory, homonuclear and heteronuclear diatomic molecules multicenter bonding in electron deficient molecules, bond strength and bond energy, percentage ionic character from dipole moment and electronegativity difference.

Unit 4 Theory of quantitative and qualitative analysis 5 Marks

Strength of acid and bases, pH, common ion effect, solubility of precipitates, solubility product. Principles of oxidimetry and reductimetry, iodimetry and iodometry.

Gravimetric analysis – its principles, precipitation, coprecipitation, postprecipitation, theory of washing.

Error in quantitative analysis

SECTION – B : ORGANIC CHEMISTRY

25 Marks : 30 Hours

Unit 1 Structure and Bonding 5 Marks

Hybridization (sp, sp2 and sp3) bond lengths and bond angles, bond energy, localized and delocalized chemical bond, van der Walls interactions, inclusion compounds, clatherates, charge transfer complexes, resonance, hyperconjugation, inductive and field effects, hydrogen bonding.

Unit 2 Mechanism of organic reactions 6 Marks

Curved arrow notation, drawing electron movement with arrows, half – headed and double – headed arrows, homolytic and heterolytic bond breaking. Types of reagents – electrophiles and nucleophiles. Types of organic reactions. Energy considerations. Reactive intermediates – carbocations, carbanions, free radicals, carbenes, arynes and nitrenes (with examples). Assigning formal charges on intermediates and other ionic species.

Methods of determination of reaction mechanism (product analysis, intermediates, isotope effects, kinetically controlled and thermodynamically controlled reactions and stereochemical studies).

Unit 3 Cycloalkanes 5 Marks

Nomenclature: monocyclic, bicyclic, tricyclic, cycloalkanes. Baeyer's strain theory and its limitations. Ring strain in small rings (cyclopropane and cyclobutane), theory of strainless rings. The case of cyclopropane ring: banana bonds.

Unit 4 Alkenes Cycloalkenes, Dienes and Alkynes 9 Marks

Methods of formation, mechanisms of dehydration of alcohols and dehydrohalogenation of alkyl halides, regioselectivity in alcohol dehydration. The Saytzeff rule, Hofmann elimination, physical properties and relative stabilities of alkenes.

Chemical reactions of alkenes – mechanisms involved in hydrogenation, electrophilic and free radical additions, Markownikoff's rule, hydroboration – oxidation, oxymercuration – reduction, Epoxidation, ozonolysis, hydration, hydroxylation and oxidation with $KMnO_4$,Polymerization of alkenes. Substitution at the allylic and vinylic position of alkenes.

Nomenclature and classification of dienes: isolated, conjugated and cumulated dienes. Structure of allenes and butadiene, methods of formation, polymerization. Chemical reactions 1,2- and 1,4- additions, Diels – Alder reaction.

Methods of formation. Chemical reactions of alkynes, acidity of alkynes. Mechanism of electrophilic and nucleophilic addition reactions, hydroboration-oxidation, metal-ammonia reductions, oxidation and polymerization.

SECTION C : PHYSICAL CHEMISTRY

25 Marks : 30 Hours

Unit 1 Gaseous State – I 6Marks

Kinetic molecular model of a gas: postulates and derivation of the kinetic gas equation; collision frequency; collision diameter; mean free path, including their temperature and pressure dependence. Barometric distribution and its use in evaluating molecular velocities (average, root mean square and most probable) and average kinetic energy, law of equipartition of energy.

Unit 2 Gaseous State – II 6Marks

Deviations from ideal gas behaviour, compressibility factor, Z, and its variation with pressure and temperature for different gases. Causes of deviation from ideal behavior, van der Waals equation of state, its derivation and application in explaining real gas behavior, mention of other equations of state (Berthelot, Dieterici); Boyle temperature. Continuity of states, critical state, relation between critical constants and van der Waals constants, law of corresponding states.

Unit 3 Liquid state 5Marks

Nature of the liquid state, intermolecular forces, Qualitative treatment of the structure of the liquid state; physical properties of liquids; vapour pressure, surface tension and coefficient of viscosity, and their determination. Effect of addition of various solutes on surface tension and viscosity. Temperature variation of viscosity and surface tension of liquids.

Unit 4 Solid state 8Marks

Nature of the solid state, law of constancy of interfacial angles, law of rational indices, Miller indices, elementary ideas of symmetry, symmetry elements and symmetry operations, qualitative idea of point and space groups, seven crystal systems and fourteen Bravais lattices; X-ray diffraction, Bragg's law, a simple account of rotating crystal method and powder pattern method.

CH-101P: INORGANIC CHEMISTRY PRACTICAL

25 Marks: 45 Hours

I. Semimicro analysis (4 radicals)

Semimicro analysis of inorganic mixtures containing four radicals/ions from the following list: Silver,lead, mercury,bismuth, copper, cadmium, arsenic, manganese, cobalt, aluminium iron, nickel, calcium,strontium,barium,magnesium,sodium, potassium,ammonium,chloride,bromide,iodide, fluoride, sulphate, sulphite,thiosulphate,chromate, phosphate, nitrate,nitrite,borate, arsenite andarsenate.

II. Quantitative analysis Volumetric Estimation (one metal)

Iodometry, dichromatometry

SEMESTER – II

CH- 202

Section A: INORGANIC CHEMISTRY

25 Marks: 30 Hours

Unit 1 Acids and Bases 6 Marks

Arrhenius concept, Bronsted-Lowry theory, electronic theory, Lux- flood theory, solvent system theory, Lewis theory of acids and bases.

Unit 2 Oxidation and Reduction 6 Marks

Electronic concept of oxidation number, concept of oxidation-reduction, oxidation- reduction potentials, factors influencing redox potential.

Unit 3 Non – aqueous solvents 6Marsks

Classification of solvents(protic,aprotic,amphiprotic),qualities of ionizing solvents,study of reactions in liquid ammonia, liquid hydrogen fluoride and liquid sulphur dioxide.

Unit 4 Chemistry of s- block elements 7Marks

Comparative studies, diagonal relationships, salient features of hydrides, salvation and complexation tendencies including their function in biosystems.

Section B: ORGANIC CHEMISTRY

25 Marks: 30 Hours

Unit 1 Stereochemistry of organic compounds 10 Marks

Concept of isomerism – elements of symmetry, molecular chirality, enantiomers, stereogenic centre, optical activity, properties of enantiomers, chiral and achiral molecules with two stereogenic centres, diastereomers, threo – and erythro – diastereomers, meso – compounds.

Relative and absolute configuration, sequence rules, D and L and R and S systems of nomenclature, Geometrical isomerism, E and Z system of nomenclature, geometrical isomerism in oximes and alicyclic compounds.

Conformation isomerism – conformational analysis of ethane and n-butane; conformations of cyclohexane derivatives. Newman projection and Sawhorse formulae, Fischer and flying wedge formulae.

Difference between configuration and conformation.

Unit 2 Arenes and aromaticity 7 Marks

Structure of benzene : molecular formula and Kekule structure, Stability and carbon-carbon bond lengths of benzene, resonance structure, MO picture,

Arornaticity: the Huckel rule, aromatic ions,

Aromatic electrophilic substitution–general pattern of the mechanism, role of σ and π-complexes and energy profile diagram.Mechanism of nitration, halogenation, sulphonation, mercuration and Friedel – Crafts reaction. Activating and deactivating substituents, orientation and ortho/para ratio.

Unit 3 Alkyl halides and aryl halides 4 Marks

Mechanisms of nucleophilic substitution reaction of alkyl halides. S_N2 and S_N1 reaction with energy profile diagrams.

Methods of formation of aryl halides, nuclear and side chain reactions. The addition – elimination and the elimination – addition mechanisms of nucleophilic aromatic substitution reactions.

Unit 4 Alcohols 4 Marks

Synthesis from carbonyl compounds,dihydricalcohols–nomenclature, methods of formation, chemical reactions of vicinal glycols, oxidative cleavage [$Pb(OAc)_4$ and HIO_4].

Trihydric alcohols- nomenclature, chemical reactions, nitration, reaction with $KHSO_4$.

Section C: PHYSICAL CHEMISTRY

25 Marks: 30 Hours

Unit 1 Solutions 6 Marks

Solutions and mixtures, miscible and immiscible liquids, types of solutions, Raoults law and Henry's laws, ideal and nonideal solutions, deviations from ideal behavior, vapour pressure of liquids and liquid mixtures, separation of completely miscible binary liquids solutions by distillation, azeotropic mixtures, solubility of partially miscible liquids (phenol – water, TEA – water and nicotine – water systems), critical solution temperature, Nernst's distribution law and its limitations.

Unit 2 Dilute Solutions 6 Marks

Dilute solutions; Colligative properties – lowering of vapour pressure.Clapeyron – Clausius equation,Thermodynamic derivation using chemical potential to derive relations between the four colligative properties [(i)relative lowering of vapour pressure, (ii)elevation of boiling point,(iii)Depression of freezing point, (iv)osmotic pressure and amount of solute. Applications in calculating molar masses of normal, dissociated and associated solutes in solution.

Unit 3 Colloids and Surface Chemistry 6 Marks

Colloidal state andcolloidal systems,characteristics of true solutions, colloidal solutions and suspensions. Classification,preparation and purification of colloidal solutions,properties of colloidal solutions: Tyndal effect,Brownian motion.Adsorption-Physisorption and chemisorption – Freundlich adsorption isotherm – Langmuir adsorption isotherm.

Unit 4 Thermodynamics – 1 7 Marks

Intensive and extensive variables; state and path functions; isolated, closed and open systems; zeroth law of thermodynamics. Firstlaw: Concept of heat,q,work, w, internal energy U and statement of first law;enthalpy, H, relation between heat capacities, calculations of q,w,U and H for reversible, irreversible and free expansion of gases(ideal and van derWaals) under isothermal and adiabatic conditions.Joule- Thomson effect and relation between Joule-Thomson coefficient and other thermodynamic parameters;inversion temperature.

CH-202P ORGANIC CHEMISTRY PRACTICAL

Marks 25: 45 Hours

1. Determination of melting point: Naphthalene 80-82°C, Benzoic acid 121.5-122°C, Urea 133.5-135°C, Succinic acid 184.5 – 185^0, trans – Cinnamic acid 133.5-135°C, cis – Cinnamic acid 58°C, Salicylic acid 157.5-158°C, Acetanilide 113.5-114°, m- Dinitrobenzene 90°, p-Dichlorobenzene 52°,Aspirin 135°.

2. Determination of boiling point: Ethanol78°,Cyclohexane81.4°,Toluene 110.6°, Benzene80°C.

3. Mixed melting point determination: Urea-Cinnamic acid mixture using of various compositions (1:4, 1:1,4:1).

4. Distillation: Simple distillation of ethanol-water mixture using water condenser. Distillation of nitrobenzene and aniline using air condenser.

5. Crystallisation: Concept of induction of crystallisation, Benzoic acid from water.

6. Decolourisation and crystallization using charcoal: Decolourisation of brown sugar (sucrose) with animal charcoal using gravity filtration.

SEMESTER – III

CH-303 Section A: INORGANIC CHEMISTRY

25 Marks: 30 Hours

Unit 1 Metallurgy 6 Marks

Minerals and ores, general principles of metallurgy, extraction of Li, K, Be, Sn, Sb, Bi, Cr and Mn.

Unit 2 Chemistry of-block elements 6Marks

Comparative studies, diagonal relationships,salient features of hydrides, oxides, oxyacids and halides, basic properties of halogens, interhalogens and polyhalogens. Applications of p-block elements (Si,Ge,Se).

Unit 3 General properties of d-block elements 6Marks

Definition, position in periodic table, Characteristic properties of d-block elements, occurrence and abundance, variable oxidation states.

Unit4 Coordination Chemistry 7 Marks

Werner's co-ordination theory and its experimental verification, types of ligands, nomenclature of coordination compounds (IUPAC), coordination number and stereochemistry of coordination compounds, isomerism of coordination compounds.

SECTION – B: ORGANIC CHEMISTRY

25 Marks: 30 Hours

Unit 1 Phenols 5 Marks

Acidic character. Comparative acidic strengths of alcohols and phenols, resonance stabilization of phenoxide ion.Reactions of phenols –electrophilic aromatic substitution, acylation and carboxylation, Mechanisms of Fries rearrangement, Claisen rearrangement, Gatterman synthesis, Hauben – Hoesch reaction and Reimer –Tiemann reaction.

Unit 2 Ethers and epoxides 5 Marks

Ethers: Methods of their formation, physical properties. Chemical reactions – cleavage and autoxidation, Ziesel's method.

Synthesis of epoxides. Acid and base – catalyzed ring opening of epoxides, orientation of epoxide ring opening, reactions of Grignard and organolithium reagents with epoxides.

Unit 3 Aldehydes and ketones 8 Marks

Synthesis of aldehydes and ketones with particular reference to the synthesis of aldehydes from acid chlorides, synthesis of aldehydes and ketones using 1, 3, - dithianes, synthesis of ketones from nitriles and from carboxylic acids. Physical properties.

Mechanism of nucleophilic additions to carbonyl group with particular emphasis on benzoin, aldol, Perkin and Knoevenagel condensations, Condensation with ammonia and its derivatives. Witting reaction, Mannich reaction.

Oxidation of aldehydes, Baeyer – Villiger oxidation of Ketones. Cannizzaro reaction, MPV reaction, Clemmensen reduction, Wolff – Kishner reduction, $LiAlH_4$ and $NaBH_4$ reductions. Halogenation of enolizable ketones.

An introduction to α, β- unsaturated aldehydes and ketones.

Unit 4 Organic compounds of Nitrogen 7 Marks

Preparation of nitroalkanes, and nitroarenes, Chemical reactions of nitroalkanes, Mechanisms of nucleophilic substitution in nitroarenes and their reductions in acidic, neutral and alkaline media, Picricacid.

Structure and nomenclature of amines, physical properties, Stereochemistry of amines, Separation of amixture of primary,secondary and tertiary amines. Structural features effecting basicity of amines. Preparation of alkyl and aryl amines (reduction of nitrocompounds, nitriles), reductive amination of aldehydic and ketonic compounds. Gabriel-phthalimide reaction, Hofmann bromamide reaction.

SECTION C: PHYSICAL CHEMISTRY

25 Marks: 30 Hours

Unit 1 Thermochemistry 6 Marks

Heats of reactions:standard states; enthalpy of formation of molecules, and ions and enthalpy of combustion and its applications; calculation of bond energy, bond dissociation energy and resonance energy from thermo chemical data, effect of temperature(Kirchoffs equations)

Unit2 Thermodynamics-II 6Marks

Carnot cycle and its efficiency, concept of entropy; thermodynamic scale of temperature,statement of the second law of thermodyamics;Calculation of entropy change for reversible and irreversible processes. Free Energy Functions and Gibbs and Helmholtz equation.

Unit3 Chemical equilibrium 7 Marks

Criteria of thermodynamic equilibrium, chemical equilibria in ideal gases, concept of fugacity.Thermodynamic derivation o f relation b e t w e e n G i b b s free energy of reaction. Equilibrium constants and their quantitative dependence on temperature, pressure and concentration. Free energy of mixingand spontaneity; thermodynamic derivation of relations between the various equilibriumconstants Kp,Kc and Kx. Le Chatelier principle.

Unit4 ChemicalKinetics – I 6Marks

Orde rand molecularity of a reaction, rate laws in terms of the advancement of a reaction, differential and integrated form of rate expressions up to.second order reactions, Zero order reactions and examples -half life period with examples, effect oftemperature on the rate of reactions - Arrehenius equation and concept of energy of activation. Experimental methods of the determination of rate laws.

CH-303P PHYSICAL CHEMISTRY PRACTICAL

25 Marks: 45 Hours

1.**Surface tension measurements** (use of organic solvents excluded).
Determine the surface tension by(i)drop number (ii)drop-weight method.

2.**Viscosity measurement:**
(a) Viscosity measurement of given liquids using Ostwald's viscometer (at room temperature)
(b)Study the effect of variation of viscosity of an aqueous solution with the concentration of solute.

3.**pHmeasurements**
a)Measurement of pH of different solutions using pH-meter.
b)Preparation of buffer solutions
(i)Sodium-acetate-acetic acid
(ii)Ammonium chloride-ammoniumhydroxide.
Measurement of the pH of buffer solutions and comparison of the values with theoretical values.
c)pH-metric titrations of
(i)strong acid and strong base
(ii)weak acid and strong base
Any other experiment carried out in the class.

SEMESTER –IV

CH-404 Section A:INORGANICCHEMISTRY

25 Marks: 30 Hours

Unit 1 Chemistry of Lanthanides 6Marks

Position of lanthanides in the periodic table, general properties of lanthanides, electronic structure, oxidation states,ionic radii and lanthanide contraction,consequences of lanthanide contraction, complex formation, uses of lanthanides and their compounds.

Unit2 Chemistry of Actinides 6Marks

Position of actinides in the periodic table, general properties of actinides, identification and nuclear synthesis of trans-uranium elements, separation of Np, Pu and Am fromU,similarities between the later actinides and later lanthanides.

Unit3 Chemistry of noble gases 6 Marks

Position in the periodic table,principles of isolation, chemical properties, bonding and stereochemistry of xenon compounds,uses of noble gases.

Unit 4 Hard and soft acids and bases 7 Marks

Classification of acids and bases as hard and soft, Pearson's concept, acid – base strength and hardness and softness. Symbiosis, theoretical basis of hardness and softness, electronegativity and hardness and softness.

Section B: ORGANIC CHEMISTRY

25 Marks: 30 Hours

Unit1 Carboxylicacids 6 Marks

Acidity of carboxylic acids,effects of substituents on acid strength. Reactions of carboxylic acids.Hell-Volhard-Zelinsky reaction. Synthesis of acid chlorides, esters and amides.Reduction of carboxylic acids. Mechanism of decarboxylation.

Hydroxy acids: malic,tartaric and citric acids.

Unit 2 Carboxylic acid derivatives 6 Marks

Relative stability of acyl derivatives.Physical properties, interconversion of acid derivatives by nucleophilic acyl substitution.

Preparation of carboxylic acid derivatives, chemical reactions. Mechanisms of esterification and hydrolysis (acidic and basic).

Unit3 Organometalliccompounds 6 Marks

Organomagnesium compounds: the Grignard reagents –formation, structure and chemical reactions.

Organozinc compounds:formation and chemicalreactions.

Organolithium compounds:formation and chemical reactions.

Unit 4 Polymers 7 Marks

Natural and synthetic, mechanism of polymerization, condensation and addition polymers, Synthetic plastics, thermosetting and thermoplastic. Urea – formaldehyde, phenol – formaldehyde plastics. Teflon, polystyrene and polyurethanes, natural and synthetic rubbers, synthetic fibres, acrylics, nylon – 6 and nylon – 66, terylene, elementary of fibre making, blended fibres.

Section C: PHYSICAL CHEMISTRY

25 Marks: 30 Hours

Unit 1 Catalysis 6 Marks

Types of catalyst, specificity and selectivity, mechanisms of catalyzed reaction at solid surfaces; effect of particle size and efficiency of the catalysts. Enzyme catalysis, Michaelis – Menten mechanism, acid – base catalysis. Theory of catalysis – adsorption and intermediate compound formation.

Unit 2 Ionic equilibria – I 7 Marks

Electrolytes and non – electrolytes, strong, moderate and weak electrolytes, ionization and ionization constant, factors affecting degree of ionization, ionic product of water. Calculation of pH of dilute solutions of weak acids and bases, common ion effect; dissociation constants of mono – and di – protic acids. Salt hydrolysis and pH for different salts. Buffer solutions; derivation of Henderson equation and its applications; buffer capacity, buffer range, buffer action and applications of buffers in analytical chemistry and biochemical processes in the human body.

Unit 3 Ionic equilibria – II 6 Marks

Solubility and solubility product of sparingly soluble salts – applications of solubility product principle. Qualitative treatment of acid – base titration curves. Theory of acid – base indicators; selection of indicators and their limitations.

Unit 4 Phase equilibria I 6 Marks

Phases, components and degrees of freedom, Gibbs Phase Rule (no derivation) for non – reactive and reactive systems; Application to one component systems – water, carbon dioxide and sulphur with applications.

CH-404P ANALYTICAL CHEMISTRY PRACTICAL

25 Marks: 45 Hours

1. To determine Hardness of water using EDTA
2. To estimate nickel using DMG
3. To estimate calcium content in chalk as calcium oxalate by permanganometry
4. To estimate reducing sugar by titration with standard Fehlings solution / Benidict’s

solution.
5. To determine the equivalent weight of the given acid sample by direct titration method with alkali
6. To determine the Saponification value of the given fat or oil sample.
7. To estimate protein in the given sample by Folin Lowry method / biuret method.
8. To estimate a reducing sugar by colorimetric method.
9. To determine the concentration of a given sample by using Lambert – Beer's law.
10. To determine the amount of iodine from a given sample (salt) by titration method.

SEMESTER – V

CH-505 INORGANIC CHEMISTRY

67 Marks: 90 Hours

Unit 1 Nuclear Chemistry and Radioactivity 7 Marks

Discovery of radioactivity, nature of radiations, separation of isotopes, binding energy, mass defect, half – life, group displacement law, artificial transmutation, artificial radioactivity. Nuclear binding energy and packing fraction. Thermonuclear reactions, radioactive tracer techniques and their applications.

Unit 2 Chemistry of compounds of non – transition elements 8 Marks

Comparative studies of s – p – block elements. Preparation and properties of bleaching powder, Portland cement and borazole. Study of Solid CO_2 and carboneous fuel (solid, liquid and gaseous). Oxides and oxyacids of phosphorous, oxides and hydrides of halogens. Chemical reactivity of Chalcogens (halides, oxyacids and peroxyacids of sulphur).

Unit 3 Chemistry of second and third transition element series 11 Marks

General characteristics, comparative treatment with their 3d-analogues (ionic radii, oxidation states, magnetic behavior, spectral properties and stereochemistry).

Vertical group and horizontal group relationship of 3d, 4d and 5d elements, oxides and halides of scandium, titanium, vanadium, chromium, manganese, iron, cobalt, nickel, copper and zinc gouts. Role of transition elements in biology.

Unit 4 Alloy and intermetallic compounds 6 Marks

Effect of alloying, types of alloys, rules for the formation of alloys, intermetallic compounds.

Unit 5 UV-visible spectroscopy 9 Marks

Fundamental laws of photochemistry (Lambert-Beer's law), molar absorptivity, energy levels of electron transition of n π* and π π* presentation of electronic spectra, application to characterization of groups like conjugated dienes, carbonyls and α, β-unsaturated carbonyl compounds, and inorganic compounds. Elementary ideas on instrumentation and sample handling.

Unit 6 Infrared Spectroscopy 9 Marks

Unit of frequency, wavelength and wavenumber, molecular vibrations - fundamental, overtone, combination tone. Fermi resonance, stretching and bending. Factors influencing vibrational frequencies (elementary treatment only), application to characterization of groups like C=N, C=O, C=C, COOR, N-H and $CONH_2$. Elementary ideas on instrumentation and sample handling.

Unit 7 Thermodynamic and kinetic aspects of metal complexes 5 Marks

A brief outline of thermodynamic stability of metal complexes and factors affecting the stability, substitution reactions of square planar complexes.

Unit 8 Environmental Chemistry 12 Marks

Environmental segment, atmosphere, composition of atmosphere, atmospheric structure, reactions in atmosphere, oxidation of sulphur dioxide, photochemical smog, oxidation of organic compounds, radionuclides in environment. Water pollution, nature of pollutants, treatment of water. Toxic chemicals in environment, biochemical effects of mercury, cadmium, lead and pesticides, control and treatment of the above trace elements, solid waste pollution, treatment and disposal.

CH-506 ORGANIC CHEMISTRY 67 Marks; 90 Hours

Unit 1 Carbohydrates 11 Marks

Classification and nomenclature, Monosaccharides, mechanism of osazone formation, constitution of glucose and fructose, chain lengthening and chain shortening of aldoses. Configuration of monosaccharides. Formation of glycosides, ethers and esters.

Determination of ring size of monosaccharides. Cyclic structure of D(+)- glucose. Mechanism of mutarotation.

Structures of ribose and deoxyribose.

An introduction to disaccharides (maltose, sucrose and lactose) and polysaccharides (starch and cellulose) without involving structure determination.

Unit 2 Amino acids, Peptides and Proteins 7 Marks

Classification, structure and stereochemistry of amino acids. Acid-base behaviour, isoelectric point and electrophoresis: Preparation and reactions of -amino acids.

Classification of proteins, Peptide structure determination, Classical Levels of protein structure. Protein denaturation/renaturation.

Unit 3 Nucleic acids 5 Marks

Nucleic acids: Introduction. Constituents of nucleic acids. Ribonucleosides and ribonucleotides. The double helical structure of DNA.

Unit 4 Fats, Oils, detergents 6 Marks

Natural fats, edible and industrial oils of vegetables origin, common fatty acids, glycerides, hydrogenation of unsaturated oils. Saponification value, iodine value, acid value. Soaps, synthetic detergents, alkyl and aryl sulphonates.

Unit 5 Pericyclic reactions 9 Marks

Definition and classification, electrocyclic reactions (thermal and photo chemical) involving 4 and 6 π- electrons and corresponding cyclo reversion reaction,cycloaddition reactions, FMO approach, Diels-Alder Reaction, photochemical [2+ 2] reactions

Unit 6 Synthetic dyes 5 Marks

Colour and constitution (electronic concept). Classification of dyes. Chemistry and synthesis of Methyl orange, Congo red, Malachite green, Crystal violet, Phenolphthalein, Fluorescein, Alizarin and Indigo.

Unit 7 Steroids 7Marks

Occurrence, nomenclature, basic skeleton, Diel's hydrocarbon and stereochemistry. Isolation, structure determination and synthesis of Cholesterol, Estrone. Biosynthesis of steroids.

Unit 8: Terpenoids 5 Marks

Occurrence, isolation, classification of terpenes, chemical composition, general methods of determining structure - Isoprene rule, synthesis and structure of citral and limonene.

Unit 9 Alkaloids 6 Marks

Definition, extraction and general methods of determining structure, isolation, structure and synthesis of nicotine, atrophine and cocaine.

Unit 10 Enzymes 6 Marks

Enzymes as biocatalyst, chemical nature, general characteristics and nomenclature of enzyme activity, Active sites, Vitamines (B complex group) and elements in enzyme function.

CH – 507 PHYSICAL CHEMISTRY 66 Marks; 90 Hours

Unit 1 Mathematics for Chemists 6 Marks

Uncertainty in measurement: types of uncertainties, combining uncertainties. Statistical treatment of uncertainties. Mean, standard deviation, relative error. Data reduction and the propagation of errors. Graphical and numerical data reduction, method of least squares (regression).

Unit 2 Atomic structure 6 marks

Bohr treatment of atomic structure and spectra of hydrogen like atoms, limitations of Bohr model. Black body radiation, Planck's theory - photo electric effect - Compton effect. Dual nature of matter, de Broglie's relationship, some simple examples.

Unit 3 Quantum Chemistry – I 8 Marks

Black-body radiation, Planck's radiation law, photoelectric effect, Bohr's model of hydrogen atom (no derivation and its defects), De Broglie hypothesis, Heisenberg's uncertainty principle. Quantum mechanical operators – momentum, position, energy

(Hamiltonian) operators, postulates of quantum mechanics. Expectation values of dynamical variables.

Unit 4 Photochemistry 6Marks

Grotthus-Drapers and Lambert Beer's Laws, Stark-Einstien's laws of photochemical equivalence, Quantum yield. Photolysis of ammonia, decomposition of Hydrogeniodide and Hydrogenchlorine reactions, Phothsynthesis. Phosophorescence, Fluorescence, Chemiluminescence and photosensitisation - definitions with examples.

Unit 5 Energetics 8 Marks

Gibbs-Helmholtz equation; Maxwell relations; thermodynamic equation of state. Systems of variable compositions, Partial molar quantities, dependence of thermodynamic parameters on composition; Gibbs-Duhem equation, chemical potential of ideal mixtures, change in thermodynamic functions in mixing of ideal gases. Nernst heat theorem, Third Law: Statement of third law, calculation of absolute entropy of molecules.

Unit 6 Specific heats of solids 6 Marks

The law of Dulong and Petit, atomic and molar heat capacities, Kopp's law, classical derivation of heat capacity, quantum theory of specific heats- Einstein equation of heat capacity of solids, Debye's equation, Debye's T law and characteristic temperatures of solids

Unit 7 Statistical Thermodynamics – I 6 Marks

Purpose of statistical thermodynamics, probability of distribution, law of multiplication of probabilities, law of addition of probabilities, Sterling approximation, concept of ensembles, canonical ensemble, microcanonical ensemble and grandcanonical ensemble.

Unit 8 Interaction of molecules with electromagnetic radiations 6 Marks

Electromagnetic radiation, wave length, wave number and frequency with their units, the electromagnetic spectrum with wave lengths and frequency, absorption of electromagnetic radiation by molecules, elementary idea of different spectroscopic techniques and the information obtainable from each.

Unit 9 Macromolecules 6 Marks

Classification of polymers - natural and synthetic - rubber, cellulose, starch, wool, silk - synthetic rubber, polyalkenes, acrylics, polyamides, polyesters, PVC polyurethane starting materials and uses only. Number average molecular weight and weight average molecular weight. Special properties of polymers.

Unit 10 Conductance 9 Marks

Metallic and electrolytic conductors - specific, equivalent and molar conductance - measurement of conductance - variation of Conductance with dilution for strong and weak electrolytes (qualitative explanation) - Transport number and its determination by Hittorffs and moving boundary method -effect of temperature and concentration - ionic mobility and ionic conductance - Kohlrausch's law and its applications - salt hydrolysis and pH of a salt solution, buffer action and explanation.

CH - 508P INORGANIC AND PHYSICAL CHEMISTRY PRACTICAL

100 Marks (Inorganic: 67 marks:Physical: 33 marks)135 Hours

Inorganic Laboratory:

I. Preparation of Inorganic complexes

a. Preparation of sodium tris(oxalato)ferrate(III)

b. Preparation of Nickel Dimethylglyoxime, [Ni(DMG)2]

c. Preparation of copper tetraammine complex, [Cu(NH3)4]SO4

d. Preparation of cis and trans-bis(oxalato)diaquachromiate

II. Estimation of two constituents from a binary mixture (one volumetrically and one gravimetrically)

Estimation of the constituents from the following mixture: Iron and calcium, iron and copper, iron and manganese, copper and zinc, silver and copper, calcium and barium, calcium and lead, calcium and magnesium, copper and chloride, copper and sulphate.

III Semimicro analysis

Semimicro analyses of five radicals containing at least one rare element (V, Mo, W, etc.) Silver, lead, mercury, bismuth, copper, cadmium, arsenic, manganese, cobalt, aluminium, iron, nickel, calcium, strontium, barium, magnesium, sodium, potassium, ammonium, chloride, bromide, iodide, fluoride, sulphate, sulphite, thiosulphate, chromate, phosphate, nitrate, nitrite, borate, arsenite, and arsenate.

Physical Laboratory

(I) Study the equilibrium of the following reactions by the distribution method:

(i) I_2 in water- Kerosene/CCl_4

(ii) $I_2(aq) + I \longrightarrow I_3(aq)$

(iii) $Cu^{2}+(aq) + nNH_3 \rightarrow Cu(NH_3)n^{2}+$

(II) Perform the following potentiometric/pH-metric titrations:

(i) Strong acid with strong base (ii) weak acid with strong base and (iii) dibasic acid with strong base

(III) Potentiometric/pH-metric titration of Mohr's salt with potassium dichromate

(IV) Determination of critical solution temperature and composition of the phenol-water system and to study the effect of impurities on it.

(V) Phase equilibria: Construction of the phase diagram of (i) simple eutectic and (ii) congruently melting systems, using cooling curves and ignition tube methods.

Any other experiment carried out in the class.

SEMESTER – VI

CH-609 INORGANIC CHEMISTRY

67 marks: 90 Hours

Unit 1 Bonding in coordination compounds 14 Marks

Theory of co-ordination bond, Effective atomic number rule, Valence bond theory and its limitations. Crystal field theory. Splitting of d-orbitals in different stereochemistries octahedral, tetrahedral and square planner complexes. Factors that influence complex formation, stability constants.

Unit 2 Magnetic properties of transition metal complexes 8 Marks

Types of magnetic behaviour, methods of determining magnetic susceptibility, spin only, formula, L-S coupling, and applications of magnetic moment data in 3d transition metal complexes.

Unit 3 Inorganic polymers 7 Marks

Silicates and their classifications and structures, phophazenes as inorganic polymers, structure and bonding in triphosphazenes, zeolites and molecular sieves.

Unit 4 Thermoanalytical methods 9 Marks

Thermogravimitric (TGA) and Differential Thermal Analysis (DTA), Differential Scanning Calorimetry (DSC) - Basic principles, Instrumentation, Factors affecting to thermoanalytical techniques, Application in soils, organic and inorganic compounds, analytical chemistry.

Unit 5 Organometallic Chemistry 9 Marks

Definition, nomenclature and classification of organometallic compounds. 18 electron rule, counting of electrons in compounds; bonding and structure of CO, NO and N_2 compounds.

Unit 6 Bioinorganic Chemistry 9 Marks

Essential and non essential trace elements in biological processes, metalloporphyrins with special reference to haemoglobin and myoglobin. Biological role of alkali and alkaline earth metal ions with special reference to Na^+, K^+ and Ca^{2+}, nitrogen fixation, chlorophyll.

Unit 7 Inorganic rings and cages 5Marks

Boron hydrides, diborane and higher boranes, borazine, tetrasulphur, tetranitride, synthesis, structure and their properties.

Unit 8 Non-stoichiometric compounds 6 Marks

Radius ratio rules, classification of ionic structures, layer structures, lattice energy, Born-Harber cycle, non-stoichiometric defects and stoichiometric defects, semiconductor and transistors, photovoltaic cells.

CH - 609 ORGANIC CHEMISTRY

66 Marks: 90 Hours

Unit 1 Organosulphur compounds 5 Marks

Nomenclature, structural features, Methods of formation and chemical reactions of thiols, thioethers, sulphonic acids, sulphonamides.

Unit 2 Elimination reactions 7 Marks

Elimination Reaction, -elimination, -elimination, The E2, E1 and El cb mechanisms, orientation effects in Elimination Reactions, stereochemistry of E2, Elimination Reactions, elimination Vs substitution, factors affecting the elimination and substitution reactions.

Unit 3 Organic synthesis via enolates 7 Marks

Acidity of -hydrogens, alkylation of diethyl malonate and ethyl acetoacetate. Synthesis of ethyl acetoacetate: the Claisen condensation Keto-enol tautomerism of ethyl acetoacetate.

Alkylation of 1, 3-dithianes. Alkylation and acylation of enamines

Unit 4 Heterocyclic compounds 10 Marks

Introduction: Molecular orbital picture and aromatic characteristics of pyrrole, furan, thiophene and pyridine. Methods of synthesis and chemical reactions with particular emphasis on the mechanism of electrophilic substitution. Mechanism of nucleophilic substitution reactions in pyridine derivatives. Comparison of basicity of pyridine, piperidine and pyrrole.

Introduction to condensed five and six-membered heterocycles. Preparation and reactions of indole, quinoline and isoquinoline with special reference to Fisher indole synthesis, Skraup synthesis and Bischler-Napieralski synthesis. Mechanism of electrophilic substitution reactions of indole, quinoline and isoquinoline.

Unit 5 Medicinal chemistry 7 Marks

Drugs and antibiotics - synthesis and structure of the following

Sulphadrugs - Sulphadiazine - sulphaguanidine

Analgesics - aspirin, phenacetin

Antimalarials - Plasmoquin, chloroquine

Antibiotics – chloramphenicol

Unit 6 Chromatography 5 Marks

Principles and application of chromatography- column, thin layer, paper, preparatory thin layer, gas chromatography, elementary ideas of instrumentation of gas chromatography.

Unit 7 Mass spectroscopy 7 Marks

Basic principle, basic compounds of double focusing instruments, molecular ions, fragmentation of molecular ions, basic rules of fragmentation, fragmentation by -bond

rupture in alkane groups, -bond rupture near functional groups, study of the nature of fragmentation and presentation of mass spectra of 2-methylpentane, cyclohexane.

Unit 8 Nuclear Magnetic Resonance Spectroscopy 8 Marks

Qualitative and conceptual treatment of the nmr phenomenon, precessional frequency, energy trasition, theory of reasonance, chemical shift, magnetically nonequivalent protons, shielding and deshielding, spin coupling, analysis of AX type spectra like, (trans-cinnamic acid, 1, 1, 2-trichloro ethane, ethyl bromide, elementary ideas on instrumentation and sample handling.

Unit 9 Electron Paramagnetic Resonance Spectroscopy 5 Marks

Elementary principle of epr., g values hyperfine splitting, epr spectra of C_6H_6(.) and CH_3 CHOCH $_2CH_3$ and their analysis.

Unit 10 Green Chemistry 5 Marks

Principles and applications of green chemistry. Introduction,advantages and disadvantages. Applications in organic synthesis, principles of ultrasound and microwave assisted organic reactions, reactions in ionic liquids.

CH -610 PHYSICAL CHEMISTRY

67 Marks; 90 Hours

Unit 1 Computer Applications in Chemistry 6 Marks

Introduction to computers and its application in chemistry: - introduction to computers - characteristics of a computer - types of computers - block diagram of a digital computer. Algorithm - Flow chart -, Applications of computer in chemistry (only selected programs) determination of molarity, normality and molality of solutions - calculation of pH.

Unit 2 Quantum Chemistry – II 7 Marks

Schrodinger wave equation in Cartesian co-ordinates) and its importance, wave function and its physical interpretations, Schrodinger equation for a free particle moving in one dimensional box and its solutions, probability distribution of electrons - radial probability distribution curves.

Unit 3 Spectroscopy 8 Marks

Rotational spectra of diatomic molecules:

Rigid rotor, moment of inertia, energy levels, selection rules, nature of spectrum, determination of bond length. Effect of isotopic substitution on the rotational spectra.

Vibrational spectra of diatomic molecules:

Harmonic oscillator: energy levels, selection rules,' nature of spectrum, determination of force constant. Anharmonic oscillator: energy levels, selection rules, nature of spectrum, fundamental band, overtones.

Raman Spectroscopy: Raman Effect, Raman scattering -Stokes lines and Anti-Stokes' lines. Raman shift.

Unit 4 Symmetry and Point groups 6Marks

Symmetry operations - products of symmetry operations of various point groups with examples, group multiplication table (C_2v, C_3v)

Unit 5 Electrochemistry I 6 Marks

Chemical cells, reversible and irreversible cells with examples. Electromotive force of a cell and its measurement, Nernst equation; Standard electrode potential and its application to different kinds of half-cells. EMF in determination of (i) free energy, enthalpy and entropy of a cell reaction, (ii) equilibrium constants, and (iii) pH values, using hydrogen, quinone-hydroquinone, glass electrodes.

Unit 6 Electrochemistry II 7 Marks

Concentration cells with and without transference, liquid junction potential, decomposition potential, electrolytic polarization, overvoltage; determination of activity coefficients and transference numbers. Qualitative discussion of potentiometric titrations (acid-base, redox, precipitation)

Theory of strong electrolytes - Debye - Huckel - Onsager theory (without detailed treatment) - verification of Onsager equation -- Wein effect and Debye - Falkenhangen effect - ionic strength - activity and activity coefficients of strong electrolytes and thelimiting equation.

Unit 7 Statistical Thermodynamics – II 6 Marks

Basic postulates of Maxwell-Boltzmann distribution law, derivation of Boltzmann distribution law, Maxwell-Boltzmann distribution law of velocities, Partition function and its physical significances, types of partition functions (derivation not included).

Unit 8 Surface Active Agents 6 Marks

Hydrophillic and hydrophobic groups, amphiphiles, classification of surfactants, surfactants in solution, miscelles and miscelles formation.

Unit 9 Chemical kinetics II 8 Marks

Collision theory and transition state theory of reaction rates, Lindemann mechanism, Steady state approximation and reaction mechanism, Kinetics of complex reactions: (i) Opposing reactions (ii) parallel reactions , (iii) consecutive reactions and (iv) chain reactions.

Unit 10 Phase equilibria II 7 Marks

Phase equilibria of two component system : solid -liquid equilibria, simple eutectic - Bi, Cd, Pb-Ag systems, desilverisation of lead., Solid solutions : compound formation with congruent melting point (Mg-Zn) and incongruent melting point, (NaCl. H_2O), ($FeCl_3$-H_2O) and $CuSO_4$-H_2O system. Freezing mixtures, acetone dry ice.

CH - 611P ORGANIC AND PHYSICAL CHEMISTRY PRACTICAL

100 Marks (Organic : 67, Physical : 33)

Organic Laboratory:

A. Qualitative Analysis

Identification of Organic Compounds; Detection of extra elements(N,S and halogens) and functional groups – phenolic, carboxylic, carbonyl, esters, amines, nitro, anilide, alcohol, halogen derivative of hydrocarbons and hydrochloride in simple organic compounds. Analysis should include detection of elements, functional group, preparation of a solid derivative. A completely dried sample of the derivative should be submitted to the examiner.

B. Organic Preparation:

(a) Acetylation of salicylic acid, aniline, glucose and hydroquinone. Benzoylation of aniline and phenol.

(b)Aliphatic electrophilic substitution; Preparation of iodoform from ethanol and acetone.

(c) Aromatic electrophilic substitution:

Nitration: Preparation of m-dinitrobenzene, p-nitroacetanilide.

Halogenation: Preparation of p-bromoacetanilide, 2, 4, 6-tribromophenol

(d)Diazotisation/ coupling : Preparation of methyl orange and methyl red.

(e) Oxidation: Preparation of benzene from toluene.

(f) Reduction: Preparation of aniline from nitrobenzene.

Physical Laboratory

1. To study changes in conductance in the following systems
   (a) strong acid-strong base
   b) weak acid-strong base and
   (c) mixture of strong acid and weak acid-strong base
2. Study the kinetics of the following reactions.
   (a) Acid hydrolysis of methyl acetate with hydrochloric acid, volumetrically or conductometrically.
   (b) Saponification of ethyl acetate.
3. Verification of Lambert-Beer's Law
4. Determination of PK (indicator) for phenolphthalein or methyl red
5. Study the formation of a complex between ferric and thiocyanate (or salicylate) ions.

Any other experiment carried out in the class.

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Inorganic Chemistry

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7. S. Sarkar, General and Inorganic Chemistry, Part II

8. G.L. Miessler and D.A. Tart, Inorganic Chemistry, Pearson Education, Low price Edition.

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4. F.A. Cotton, Chemical Applications of Group Theory, WILEY Eastern

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8. R.K. Bansal, Synthetic Approaches in Organic Chemistry, Narosha Publishing House.

9. W. Kemp and A. Bahl, Experimental Organic Chemistry, S. Chand and Co.

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